

pH Values at Specified $[\text{H}_3\text{O}^+]$

Solution	$[\text{H}_3\text{O}^+]$ (M)	pH
1.00 L of H_2O	1.00×10^{-7}	7.00
0.100 mol HCl in 1.00 L of H_2O	1.00×10^{-1}	1.00
0.0100 mol HCl in 1.00 L of H_2O	1.00×10^{-2}	2.00
0.100 mol NaCl in 1.00 L of H_2O	1.00×10^{-7}	7.00
0.0100 mol NaOH in 1.00 L of H_2O	1.00×10^{-12}	12.00
0.100 mol NaOH in 1.00 L of H_2O	1.00×10^{-13}	13.00